



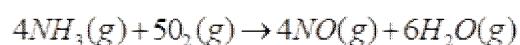
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## Test Paper -01

### Chapter 01 Chemical Reactions and Equations

#### Que. 1 Multiple choice questions (4marks)

1. The following reaction is an example of a



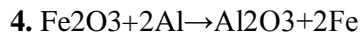
- (i) displacement reaction
  - (ii) combination reaction
  - (iii) redox reaction
  - (iv) neutralisation reaction
- (a) (i) and (iv)  
(b) (ii) and (iii)  
(c) (i) and (iii)  
(d) (iii) and (iv)

2. Electrolysis of water is a decomposition reaction. The mole ratio of hydrogen and oxygen gases liberated during electrolysis of water is

- (a) 1:1
- (b) 2:1
- (c) 4:1
- (d) 1:2

3. In the double displacement reaction between aqueous potassium iodide and aqueous lead nitrate, a yellow precipitate of lead iodide is formed. While performing the activity if lead nitrate is not available, which of the following can be used in place of lead nitrate?

- (a) Lead sulphate (insoluble)
- (b) Lead acetate
- (c) Ammonium nitrate
- (d) Potassium sulphate



The above reaction is an example of a

- (a) combination reaction
- (b) double displacement reaction
- (c) decomposition reaction
- (d) displacement reaction

**Que. 2 Answer the following questions.(any 3 out of 5) (6marks)**

1. Why should a magnesium ribbon be cleaned before burning in air?

2. is balanced chemical equation? Why should chemical equation be balanced?

3. Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

4. Write the balance equation for the following reactions .

i. Hydrogen + Chlorine  $\rightarrow$  Hydrogen chloride

ii. Barium chloride + Aluminium sulphate  $\rightarrow$  Barium sulphate + Aluminium chloride

5. What happens when a piece of

(a) zinc metal is added to copper sulphate solution?

(b) aluminium metal is added to dilute hydrochloric acid?